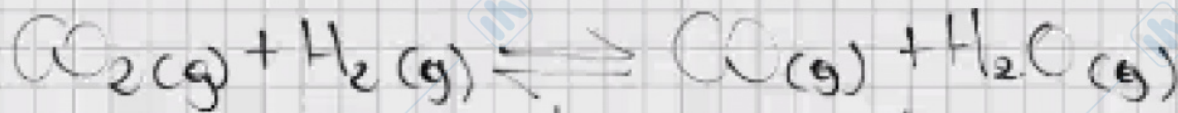


5 l



1,6 mol

1,6 mol

0 mol

0 mol

-0,6 mol

-0,6 mol

0,6 mol

0,6 mol

equilib.

1 mol

1 mol

 $K_c?$

$$M_{\text{CO}_2} = \frac{1,6 \text{ mol}}{5 \text{ l}} = 0,32 \text{ M}$$

$$M_{\text{H}_2} = \frac{1,6 \text{ mol}}{5 \text{ l}} = 0,32 \text{ M}$$

$$M_{\text{CO}} = \frac{0,6 \text{ mol}}{5,0 \text{ l}} = 0,12 \text{ M}$$

$$M_{\text{CO}_2} = \frac{1,0}{5,0} = 0,2 \text{ M}$$

Kc?

$$K_{CO_2} = \frac{1,6 \text{ mol}}{5 \ell} = 0,32 \text{ M}$$

$$K_{H_2} = \frac{1,6 \text{ mol}}{5 \ell} = 0,32 \text{ M}$$

$$K_{CO} = \frac{0,6 \text{ mol}}{5,0 \ell} = 0,12 \text{ M} \quad K_{CO_2} = \frac{1,0}{5,0} = 0,2 \text{ M}$$

$$K_{H_2O} = \frac{0,6 \text{ mol}}{5,0 \ell} = 0,12 \text{ M} \quad K_{H_2} = \frac{1,0}{5,0} = 0,2 \text{ M}$$

$$K_C = \frac{[CO][H_2O]}{[CO_2][H_2]} = \frac{0,12 \cdot 0,12}{0,2 \cdot 0,2} =$$