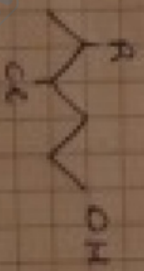


CHIMICA ORGANICA

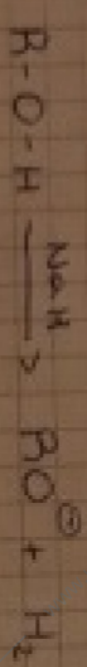
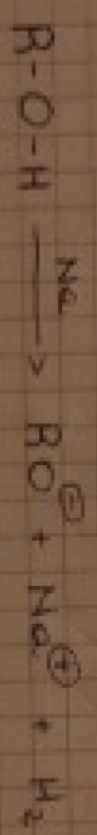
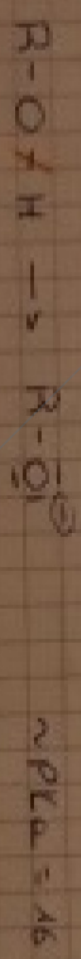
R-OH → Alcololo R-O-H

REATTIVITA' DEGLI ALCOLI

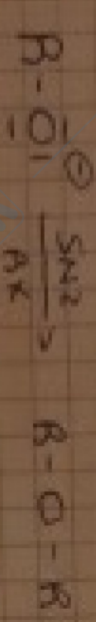


R-O-H \xrightarrow{Na} R-O⁻Na⁺ + H₂ (R-OH che i legami O-H che i legami R-O)

PRIMO CASO: DEPROTONAZIONE



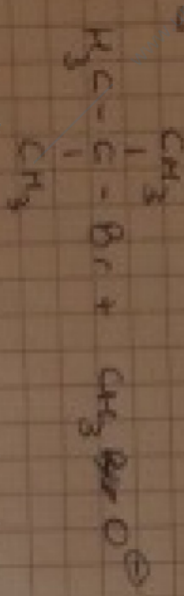
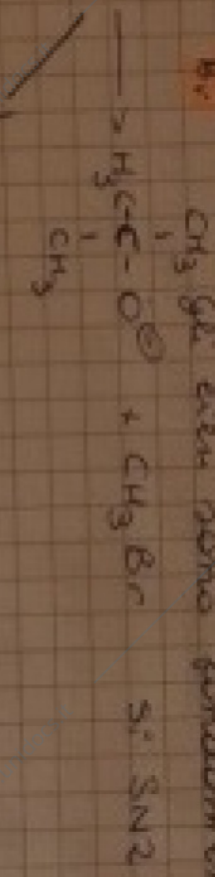
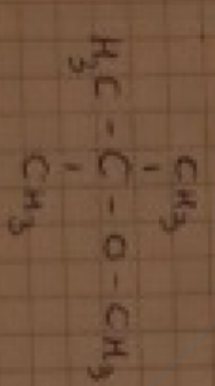
Reazione di ossidazione e' decarboxilazione ma una SN2 e formare un etere

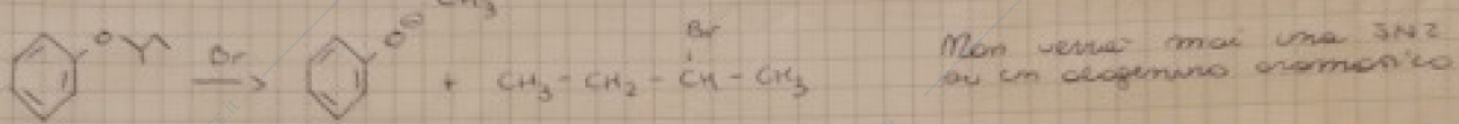
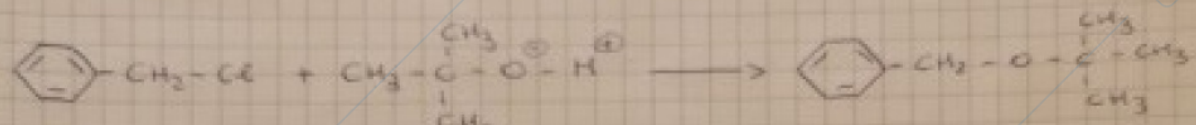
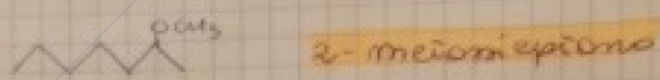
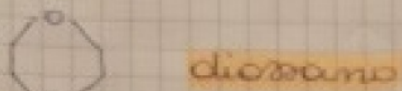
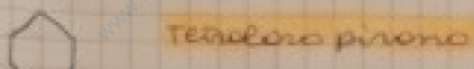
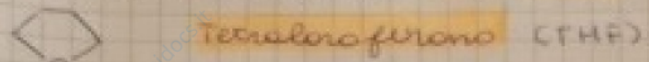
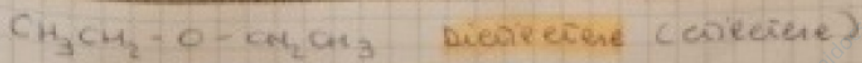


Questa reazione e' detta reazione di Williamson

RX = met, RCH₂X, R-OH-R

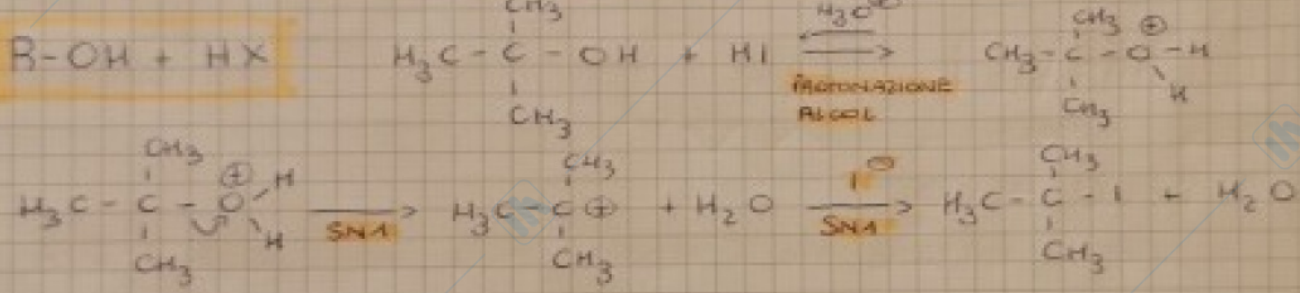
E' preferibile un metilico o un primario



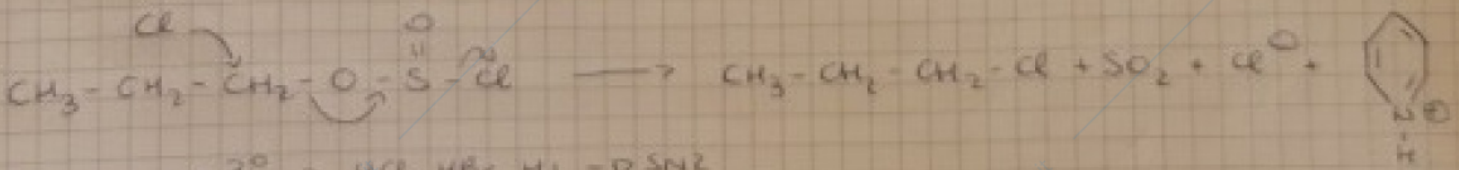
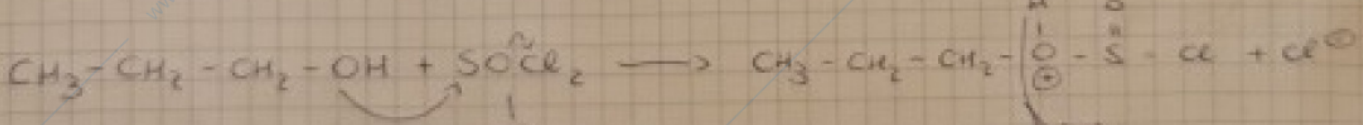
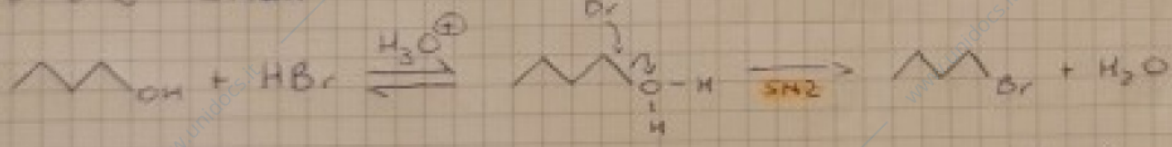


Gli alogenuri per la reazione di Williamson possono essere al massimo **secondari**

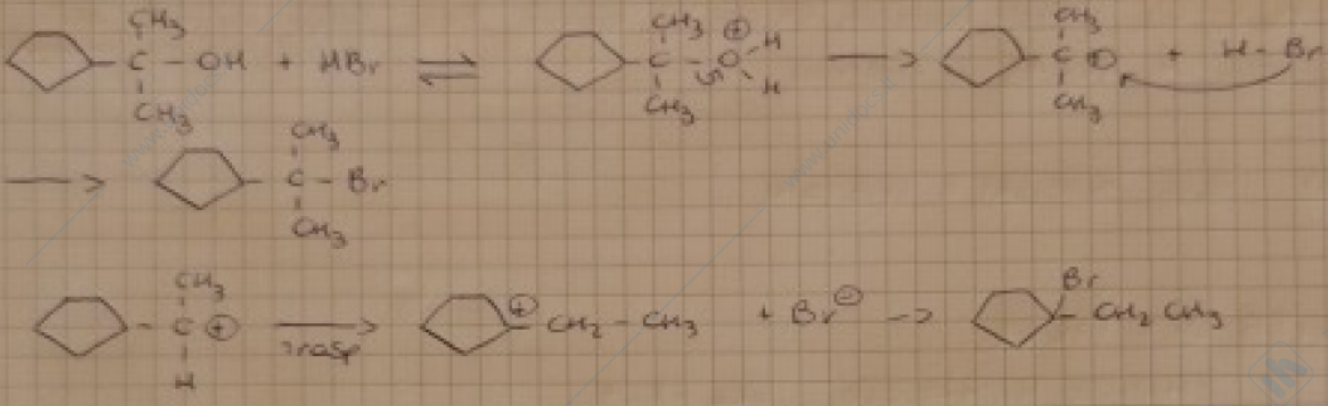
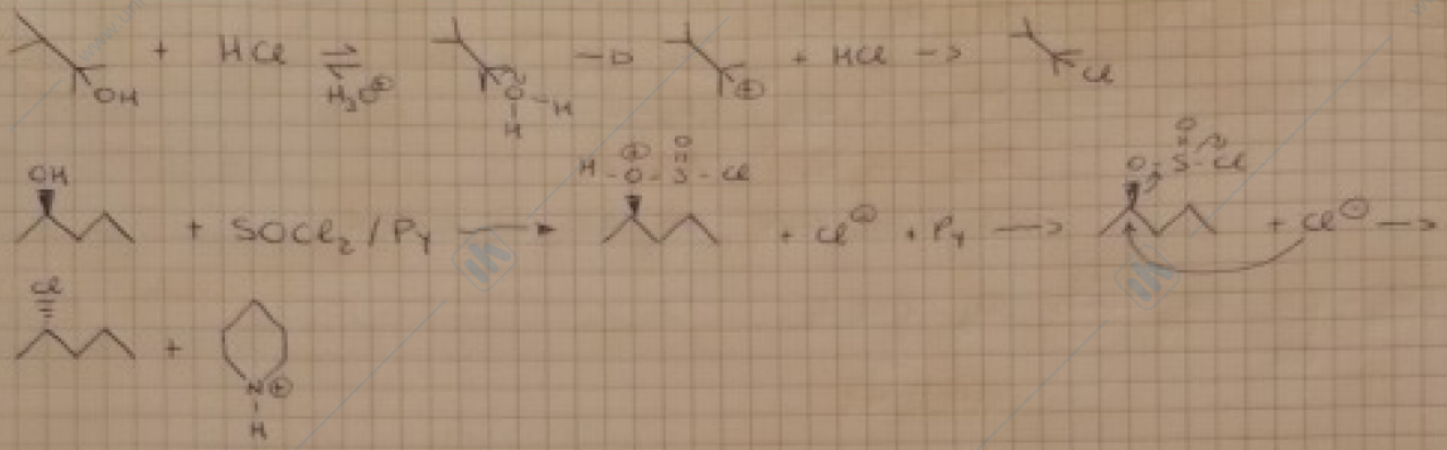
TRASFORMAZIONE DI UN ALCOL IN ALOGENURO ALKILICO



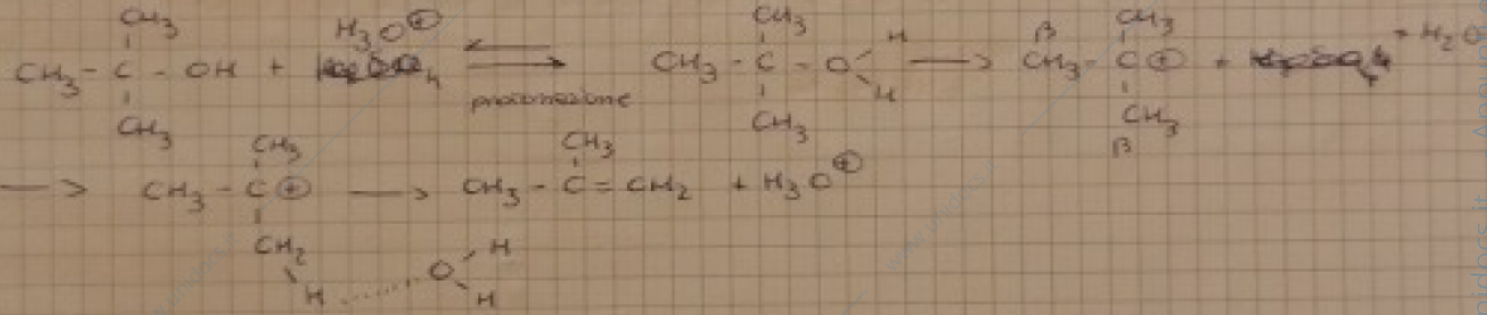
3° e 2° $\text{S}_{\text{N}}2$



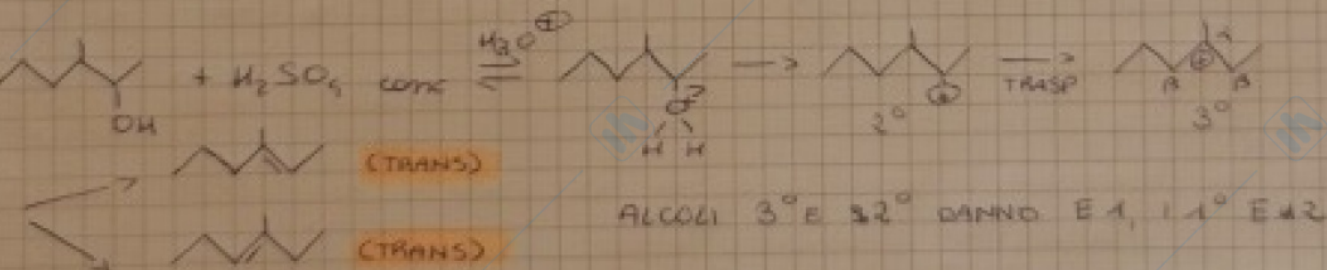
- R-OH \rightarrow 3° -> HCl, HBr, HI - $\text{S}_{\text{N}}1$
- R-OH \rightarrow 2° -> HBr, HI - $\text{S}_{\text{N}}2$
- R-OH \rightarrow 2° -> SOCl_2 + PY
- R-OH \rightarrow 1° -> SOCl_2 + PY



B-ELIMINAZIONE DEGLI ALCOL

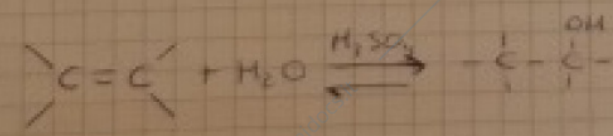


Reazione acidocatalizzata

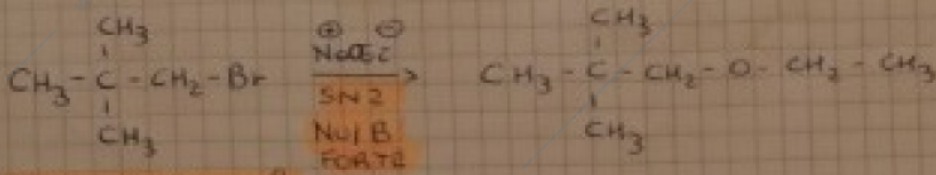


ALCOLI 3° E 2° DANNO E1, 1° E2

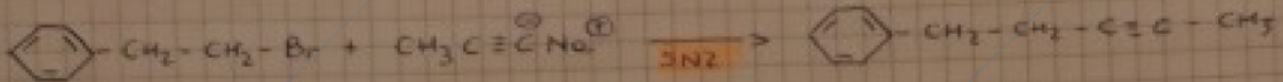
Per avere l'alkene deve lavorare in condizioni con l'acido concentrato.
 Per idratare l'alkene deve avere più acqua.



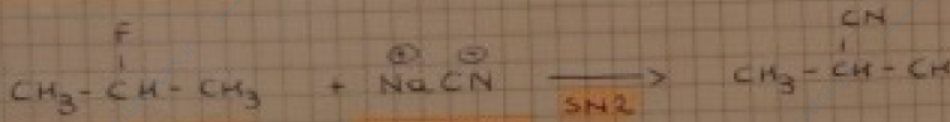
ESERCITAZIONE



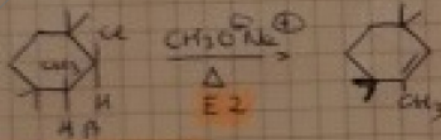
ALCOGENURO 1°
NON CI SONO H^β



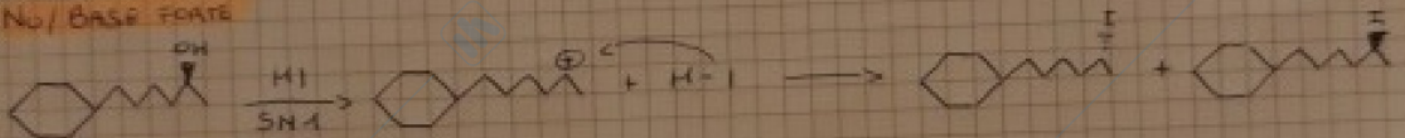
ALCOGENURO 1° NU. FORTE



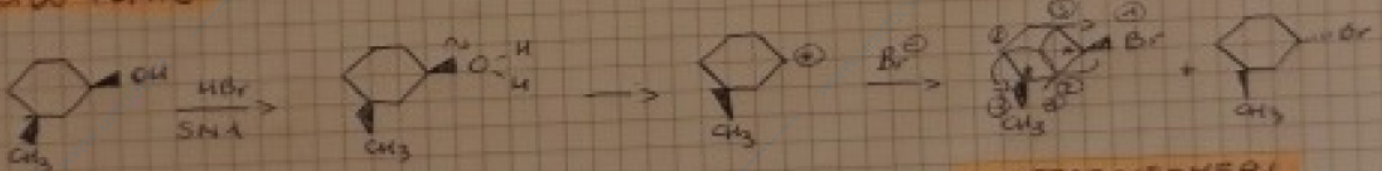
ALCOGENURO 2° NU. FORTE



ALCOGENURO 2°
NU/BASE FORTE



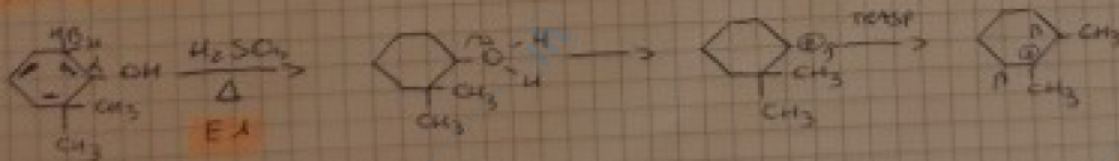
ALCOLE SECONDARIO
ACIDO FORTE



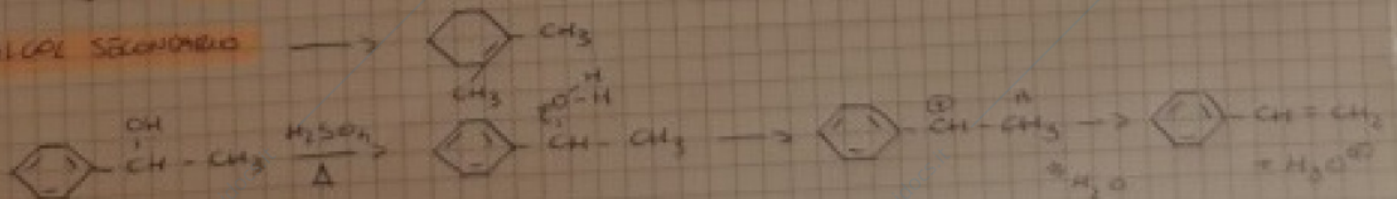
ALCOLE SECONDARIO
ACIDO FORTE

DIAStereoisomeri

(R,S) e l



ALCOLE SECONDARIO



A. SECONDARIO