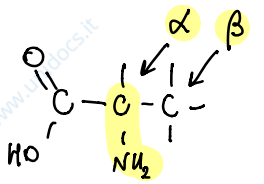


* equilibrio chimico ≠ favore di una o l'altra
 ↓
 n. portano anche atomi
 ↓
 n. portano solo elettroni

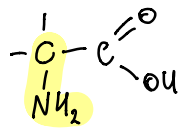
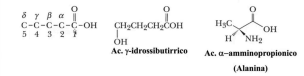
TAUTOMERIA CHETO-ENOLICA → Perché n. portano un protone e un doppio legame π



ALANINA

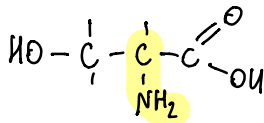
α-AMMINOACIDO

Quando si usano i nomi comuni si aggiungono spesso le lettere greche α, β, γ, δ, ε per indicare la posizione dei sostituenti

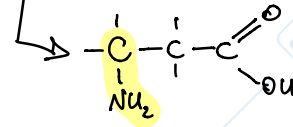


GLICINA

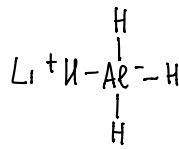
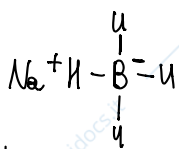
β-AMMINOACIDO



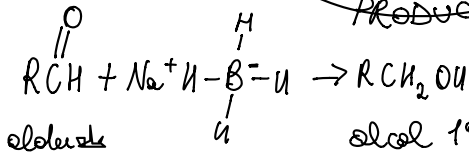
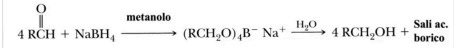
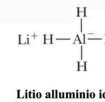
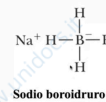
SERINA



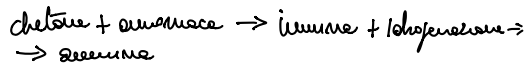
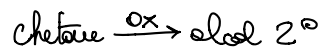
COME SI PRODUCE H⁻?



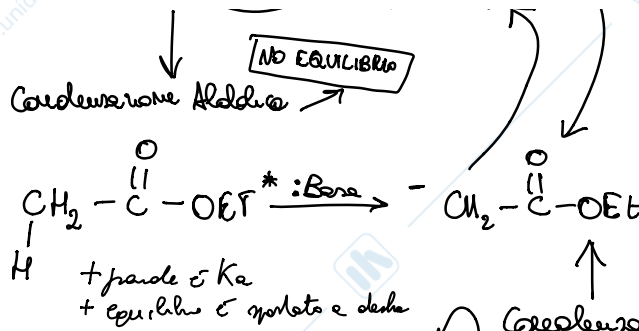
Riduzione con idruri metallici



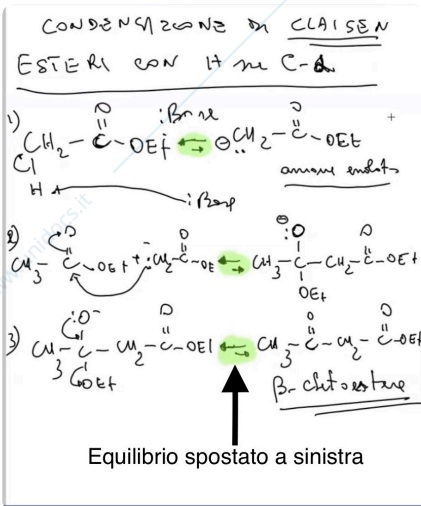
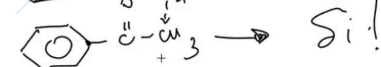
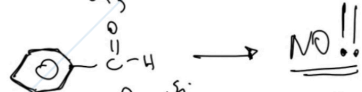
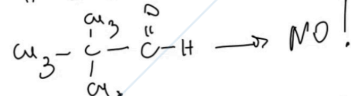
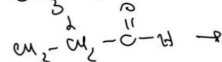
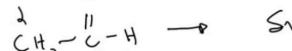
PRODUCONO



perché B e Al sono più elettropositivi di H



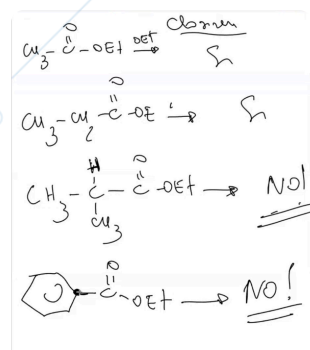
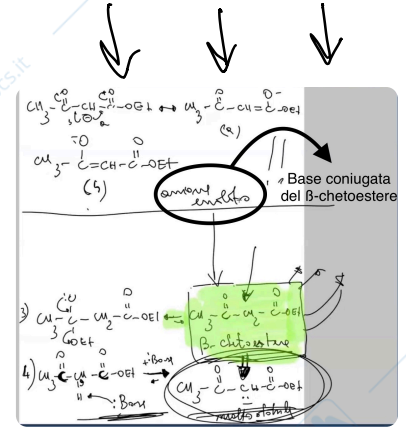
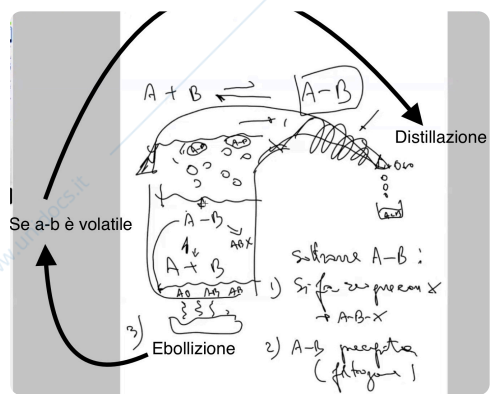
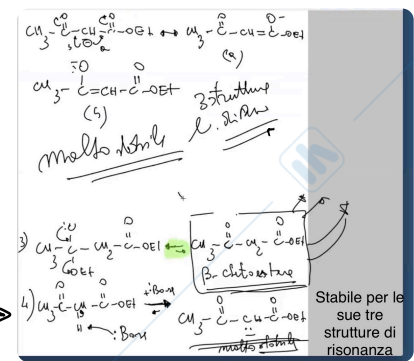
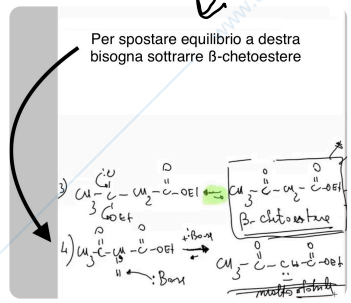
REAZIONE ALDOLICA (CONDENS. H₂O)



Condensazione di Claisen EQUILIBRIO SI

x spostare equilibrio a destra bisogna sottrarre il prodotto.

K_a acido acetico = $1,8 \times 10^{-5}$



Condensazione di Claisen → NO CH₃ / alchini (alcilati di esteri)

Condensazione aldolica → SI OH⁻